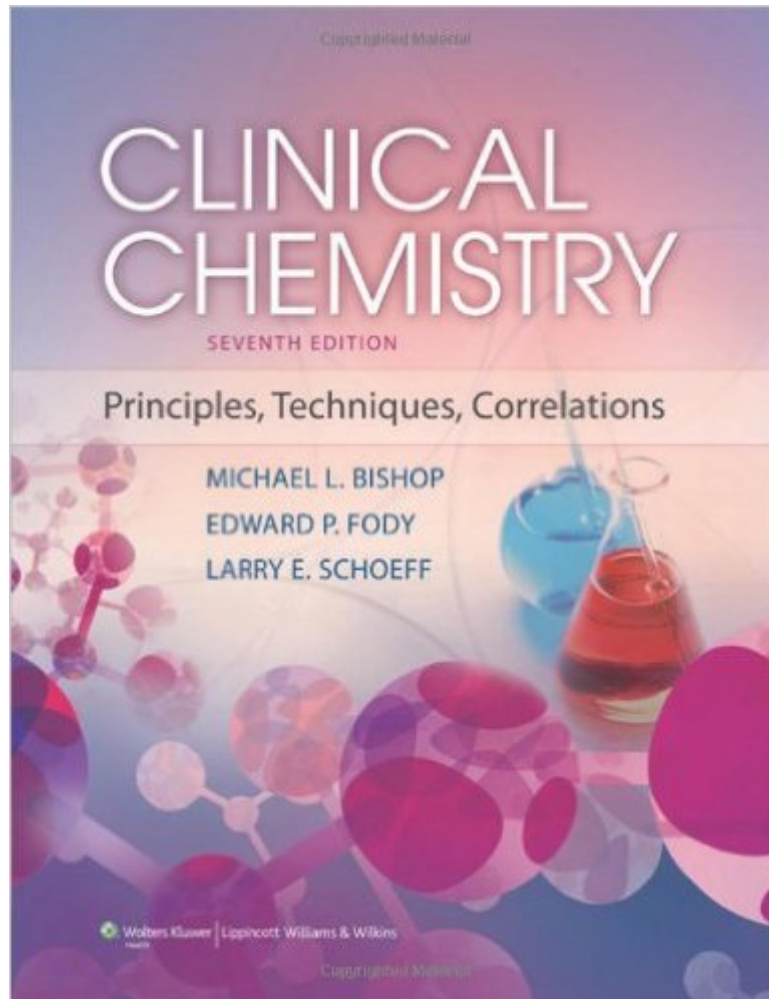


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# Clinical Chemistry: Principles, Techniques, And Correlations



## Synopsis

In its Seventh Edition, this acclaimed Clinical Chemistry continues to be the most student-friendly clinical chemistry text available. This edition not only covers the how of clinical testing but also places greater emphasis on the what, why, and when in order to help today's students fully understand the implications of the information covered, as well as the applicability of this crucial topic in practice. With clear explanations that strike just the right balance of analytic principles, techniques, and correlation of results with disease states, this edition has been fully updated with the latest information to help keep today's students at the forefront of today's science. New case studies, practice questions, and exercises provide ample opportunities to review and apply the topics covered through the text.

## Book Information

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## Customer Reviews

This book was not helpful. There is no glossary or table of contents. There is no answer key in the book to the questions at the end of each chapter. You need to go online to view the correct answers, which is very cumbersome. The book was hard to read and because of the lack of glossary, it made things very difficult to find in the book. I strongly recommend a different clinical chemistry book if there is one out there.

I am a medical lab science student and this was the required text for our clinical chemistry course. I

was not a fan. The text not written to be student friendly. It inadequately explains the concepts for which it is intended to teach. The author included unnecessary information that just muddied the water and could have been omitted. Example: In talking about the body's use and measurement of iron levels, the chapter is opened with a discussion of iron ore and its role as a naturally occurring element in the earth's crust. Save that for a geology text. Lastly, my biggest beef is failing to put the answers in the back of the book for the chapter review questions. That was very annoying to say the least!!!

I have the 5th edition of this book as well. The 7th edition is STILL missing information, and case study questions do not have relevant information in the chapter they are located, leading to internet searches to find the answers. Would be nice to have a place for answers in the back of the book like most textbooks, or on the online companion. As with pretty much all laboratory textbooks, lacking and behind in many areas of technology.

Given how my professor related the material in her lectures, anyway. Very boring book. Pretty hard to read. Very technical. Doesn't make a boring subject easier to tolerate. Glad this class is almost over, that's for sure.

The glossary leaves a bit to be desired and has to be accessed online. There is no way to access the answers to the end of chapter questions unless your professor gives them to you. This book only contains chapters and index all the other useful reference tables etc. have to be accessed online.

This book might be a good reference for people in the field, but terrible if this is required for your clinical chemistry or quantitative analysis course. It is highly granular and the main points or concepts are more difficult to parse than a regular textbook. Also many sections do not clearly answer the chapter objectives at the beginning, sometimes not touching upon them at all. There are small errors in the book that surprise me for making it through seven editions. It's written by so many authors that the writing style and quality varies with each chapter.

Good, lots of great information, but learning this material can be overwhelming. Pro -- This book comes with hardback. If you care about keeping your books hardback bound are the best. Con -- No answers to any of the chapter case studies.

As well as agreeing with most of the previous commentary complaining about the extraneous material, there are significant errors with even the simplest of calculations. Admittedly, being a chemist and having been a chemist for many years helps here, but common sense would suffice just as well. For example, in Example 1-2, it appears that, in order to produce a 10% NaOH solution, you take 100 grams of what is already a 10% solution and dilute it ! Apparently, you can produce an infinite supply of NaOH; all you have to do is dilute it ! The example before it is worse. They say to make up a 5% solution of HCl, add 5 grams of 12M HCl. Let's see what percentage that really is. Since HCl is a gas, it's not a simple matter to just add 5 grams of it to a flask; so we have to use a concentrated solution, which here is 12M. Fine so far. Specific gravity of 12M HCl is about 1.19 kg/liter. 5 grams of 12M HCl would then be:  $5 \text{ grams} (1 \text{ kg} / 1000 \text{ grams})(1 \text{ liter} / 1.19 \text{ kg}) = 0.0042 \text{ L}$ , or 4.2 mLs of 12M HCl.  $0.0042 \text{ L} (12 \text{ M} / \text{L})(36.46094 \text{ g} / \text{mole}) = 1.84 \text{ grams of HCl}$  So, adding 5 grams of 12M HCl produces a solution which is only 1.84 % HCl. Other examples in the book are just as faulty. Sure, this is picky, but that's the way it is in chemistry, and clinical chemistry in particular. How would you feel if your blood glucose level came back at 4 times what the actual value is? Yes; results have to be accurate, and the calculations in this book just aren't. If you use the approach(es) detailed in the book, check the numbers twice then ask somebody who knows what they're doing.

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